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PREVIOUS YEAR PAPER · 2020-21

OLYMPIAD · STAGE I

IOQC

Indian Olympiad Qualifier in Chemistry (Part I)

2020-21

Question paper with official answer key

32	120	60 min	31
Questions	Maximum marks	Time	Paper code

Question paper and official answer key: Indian Association of Physics Teachers (IAPT), IOQC 2020-21, paper code 31.
Solutions and notes: Zigyan. This booklet is not an IAPT publication.

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	99.95 percentile



Sulokya Bibhuty

3 YEARS CLASSROOM

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OLYMPIAD RESULTS

Above the Minimum Admissible Score - the national band for the subject

CHEMISTRY

- Sulokya Bibhuty — 3 Years Classroom
- Anay Gupta — 3 Years Classroom
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- Harshit Goyal — 4 Years Classroom
- Hardik Sharma — 3 Years Classroom
- Aarush Kumar — 1 Year Classroom

PHYSICS

- Sulokya Bibhuty — 3 Years Classroom
- Saksham Bishnoi — 4 Years Classroom

ASTRONOMY

- Sulokya Bibhuty — 3 Years Classroom

MATHEMATICS (IOQM)

- Priyanshi — 4 Years Classroom

Saksham Bishnoi qualified the Chemistry Olympiad twice and was selected for the Orientation-cum-Selection Camp at HBCSE, Mumbai.

16 selections	16 of 35	AIR 168
in JEE Advanced 2026	students of the batch selected	best JEE Advanced rank

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Given in the question paper

Useful constants

Charge of electron, $e = 1.602 \times 10^{-19} C$

Mass of electron, $m_e = 9.1 \times 10^{-31} kg$

Planck's constant, $h = 6.626 \times 10^{-34} J\,s$

Speed of light, $c = 3.0 \times 10^{8} ms^{-1}$

Avogadro constant, $N_A = 6.022 \times 10^{23} mol^{-1}$

Molar gas constant, $R = 0.082\, L\,atm\,mol^{-1}K^{-1}$

$= 8.314\, J\,mol^{-1}\,K^{-1}$

PART A-1 · ONE CORRECT OPTION

Q. 1 – 24 · +3 for a correct answer, –1 for a wrong one

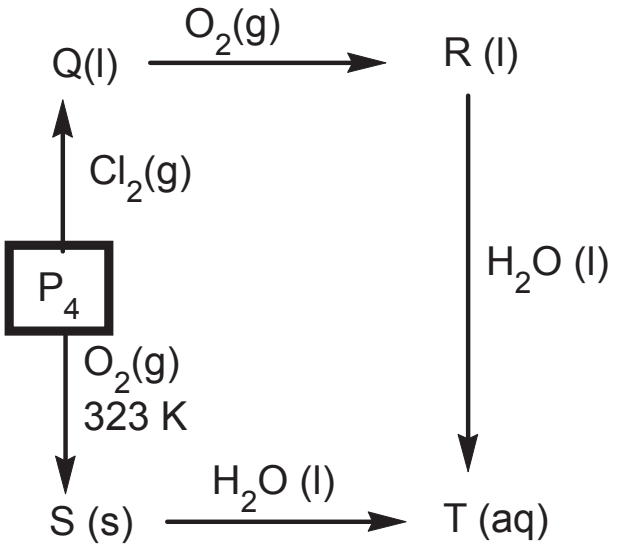
1. Reaction of ammonia with diborane gives an ionic product ($B_2H_6.2NH_3$).

The hybridization of boron in the cation and anion of this product are respectively

(A) sp^3 in both (B) sp^3 and sp^2 (C) sp^2 and sp^3 (D) sp^2 in both

2. A sequence of reactions of phosphorous (P_4) is given below

The correct set of products (Q, R, S and T) among the following is



(A) Q = PCl_3; R = $POCl_3$; S = P_2O_3; T = H_3PO_3

(B) Q = PCl_5; R = P_2O_5; S = P_4O_6; T = H_3PO_3

(C) Q = PCl_3; R = $POCl_3$; S = P_4O_{10}; T = H_3PO_4

(D) Q = PCl_5; R = P_4O_{10}; S = P_4O_{10}; T = H_3PO_4

3. In the gaseous state of $Fe(CO)_5$, the 'd' orbital that would participate in hybridization is

(A) $d_{x^2-y^2}$ (B) d_{z^2}

(C) d_{xz} (D) any one of the 'd' orbitals

4. Among the following, the *correct* statement/s about 'p' block elements is/are

I. The valence shell electronic configuration of all of them is $ns^2\ np^{1-6}$

II. Only in p block, metals, nonmetals and metalloids are present

III. Halogens have the lowest negative electron gain enthalpy in the respective periods

IV. Noble gases have no tendency to accept an electron and hence they have large negative values of electron gain enthalpy

(A) I, IV (B) II, III (C) IV only (D) II only

5. A chemical reaction is carried out at two different temperatures T_1 and T_2 ($T_2 > T_1$) and also with and without a catalyst.

The statement that is correct among the following is

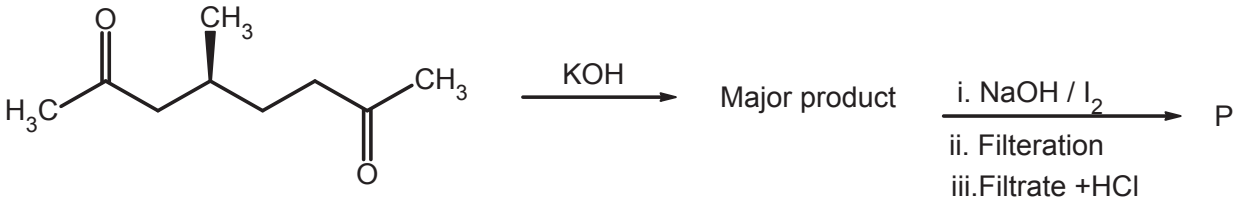
(A) Lowering in the activation energy of the reaction due to catalyst would be higher at T_2 than at T_1

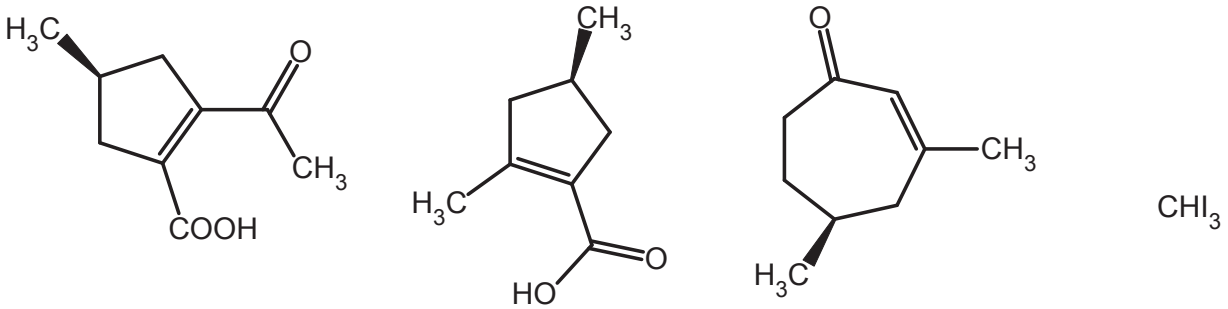
(B) Lowering in the activation energy of the reaction due to catalyst would be higher at T_1 than at T_2

(C) The factor by which the rate of the reaction is increased by the catalyst would be lower at T_2 than at T_1

(D) The factor by which the rate of the reaction is increased by the catalyst would be higher at T_2 than at T_1

6. The product **'P'** in the following sequence of reactions is





(A) (B) (C) (D)

(A) (A) (B) (B) (C) (C) (D) (D)

7. Among the following, maximum number of resonance structures is possible for

(A) PO_4^{-3} (B) SO_4^{-2} (C) CO_3^{-2} (D) MnO_4^-

8. A mixture of sodium (Na) and potassium (K) metals weighing 32 g was reacted with water and the solution obtained could be neutralized with 517.3 mL of 1.0 M H_2SO_4 (aq).

The mass of sodium that was present in the mixture is

(A) 20 g (B) 16 g (C) 10 g (D) 12 g

9. The mass ratio of steam and hydrogen is found to be 1:1.5 at equilibrium in the following reaction

$3Fe\ (s) + 4H_2O\ (g) \rightleftharpoons Fe_3O_4(s) + 4H_2\ (g)$

The value of the equilibrium constant **(K_c)** of the above reaction is

(A) 3.0×10^{-5} (B) 3.3×10^{4} (C) 3.3×10^{6} (D) 1.3×10^{3}

10. Two students did a set of experiments on ketones **'X'** and **'Y'** independently and obtained the following results.

Reaction /Experiment	**X**	**Y**
Optical rotation	Yes	Yes
Optical rotation after treatment with a base	Zero	Yes
NH_2NH_2, KOH, Heat	Formation of an optically inactive hydrocarbon C_6H_{12}	Formation of an optically inactive hydrocarbon C_6H_{12}

The ketones **'X'** and **'Y'** are respectively

(A) 2-ethylcyclobutanone and 3-ethylcyclobutanone

(B) 2-methylcyclopentanone and 3-methylcyclopentanone

(C) 3-methylcyclopentanone and 2-methylcyclopentanone

(D) 3-methyl-4-penten-2-one and 4-methyl-1-penten-3-one

11. Glycine ($C_2H_5O_2N$) is the simplest of amino acids. Molecular formula of the linear oligomer synthesized by linking *ten* glycine molecules together via a condensation reaction would be

(A) $C_{20}H_{32}O_{11}N_{10}$ (B) $C_{20}H_{68}O_{29}N_{10}$

(C) $C_{20}H_{40}O_{10}N_{10}$ (D) $C_{20}H_{50}O_{20}N_{10}$

12. If Ni^{2+} is replaced by Pt^{2+} in the complex ion $[NiCl_2Br_2]^{2-}$, which of the following would change?

I. Magnetic moment

II. Geometry

III. Geometrical isomerism

IV. Optical isomerism

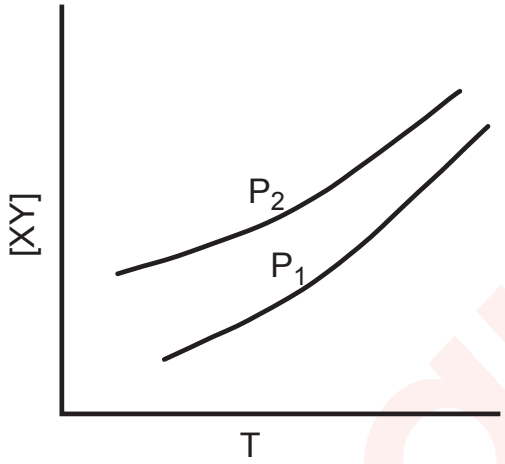
(A) I, II, III (B) II, III (C) I, II (D) II, III, IV

13. An inorganic compound **'X'** of an alkali metal on heating gives a reddish-brown gas **'Y'** and a binary solid **'Z'**. This solid is less soluble in water and its solution is basic. **'X'** does not give a positive silver nitrate test. **'X'** can be identified as

(A) KIO_3 (B) $LiNO_3$ (C) $NaNO_3$ (D) KNO_2

14. The qualitative plots given represent the yield of the product, [XY], at equilibrium in the reaction X(g) + Y(g) $\rightleftharpoons$ XY(g), as a function of temperature, at total pressures P_1 and P_2.

The reaction is



(A) endothermic and $P_1 < P_2$

(B) endothermic and $P_2 < P_1$

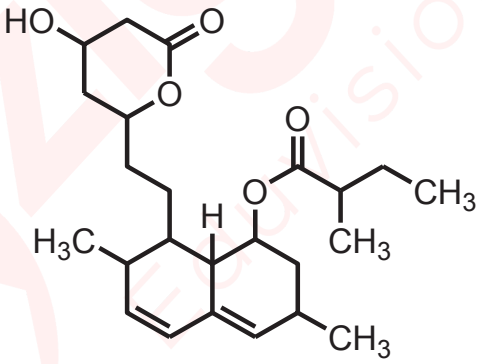
(C) exothermic and $P_1 > P_2$

(D) exothermic and $P_2 > P_1$

15. When 6.8 g of $AgNO_3$ completely reacts with H_3PO_2, metallic silver produced (g) and H_3PO_2 consumed (mole) are respectively

(A) 4.32 and 0.1 (B) 1.08 and 0.01 (C) 4.32 and 0.01 (D) 2.16 and 0.01

16. Lovastatin, a drug used to reduce the risk of cardio vascular diseases has the following structure



Lovastatin

The number of stereogenic centers present in lovastatin is

(A) 8 (B) 3 (C) 4 (D) 6

17. Among the following sets, the one in which all the molecules are non polar is

(A) XeF_4, XeO_3, XeO_4

(B) XeF_2, XeO_4, $XeOF_4$

(C) XeF_2, XeF_4, XeO_4

(D) XeF_2, XeO_3, $XeOF_4$

18. Gas phase reactions (i) and (ii) are of first and second order respectively

$2N_2O_5 \rightarrow 4NO_2 + O_2$ ..... (i)

$2NO + O_2 \rightarrow 2NO_2$..... (ii)

Under certain conditions, the rate constants (k_1, k_2) of (i) and (ii) respectively, have the same numerical value, when the concentrations of the reactants are expressed in mol/dm^3.

If the concentrations are expressed in mol/mL, the correct relationship between k_1 and k_2 is

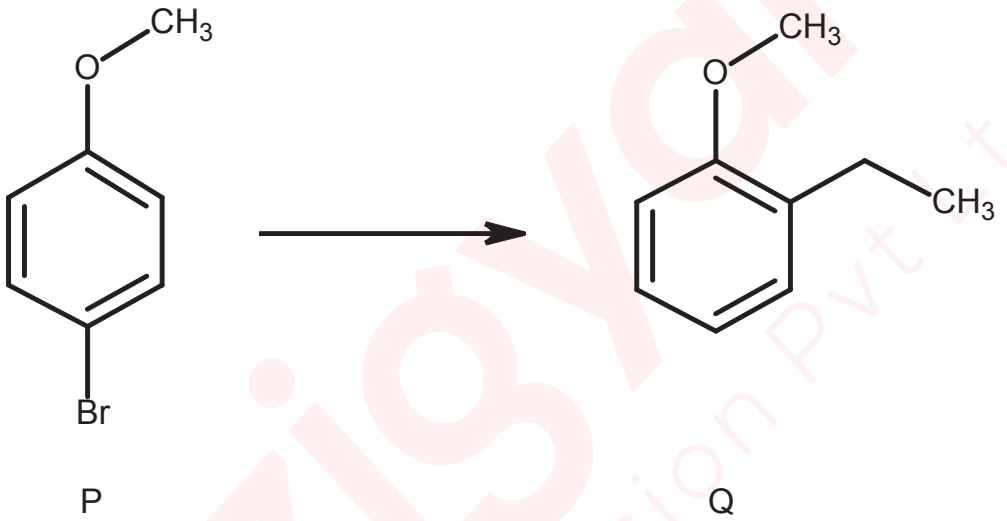
(A) $k_2 \times 10^{-3} = k_1$

(B) $k_2 \times 10^{3} = k_1$

(C) $k_1 = k_2$

(D) $k_1 \times 10^{6} = k_2$

19. The correct sequence of reactions to get **'Q'** as the *only* product from **'P'** is



P Q

(A) (i) H_2 & Pt catalyst (ii) C_2H_5Cl & $AlCl_3$

(B) (i) Mg in ether (ii) aqueous alcohol (iii) C_2H_5Cl & $AlCl_3$

(C) (i) Mg in ether (ii) C_2H_5Cl & $AlCl_3$

(D) (i) C_2H_5Cl & $AlCl_3$ (ii) Mg in ether (iii) aqueous alcohol

20. The Galvanic cell can be represented as Zn / Zn^{2+} (0.1M) // Cu^{2+} (0.1M)/Cu.

Among the following, the cell that can produce an EMF more than that of the Galvanic cell is

(E^0 of Zn^{2+}/Zn and Cu^{2+}/Cu are −0.763V and 0.337V respectively)

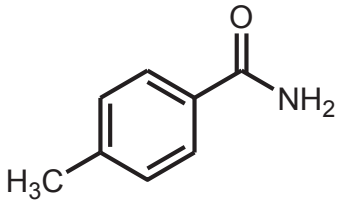
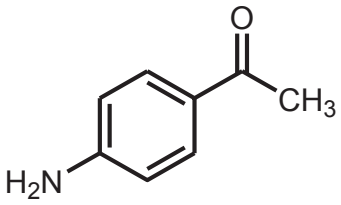
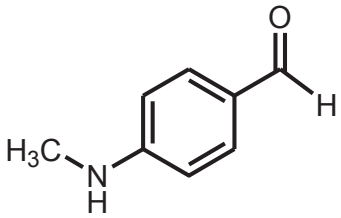
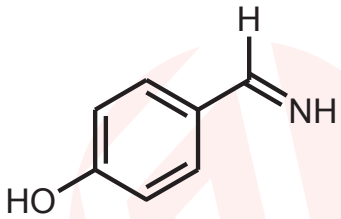
(A) Zn / Zn^{2+} (0.1M) // Cu^{2+} (0.01M)/Cu

(B) Zn / Zn^{2+} (1M) // Cu^{2+} (0.01M)/Cu

(C) Zn / Zn^{2+} (0.01M) // Cu^{2+} (1M)/Cu

(D) Zn / Zn^{2+} (0.01M) // Cu^{2+} (0.01M)/Cu

21. The correct match of the molecules in column I and reactions in column II is

Column I	Column II
i) 	(L) Coloration with $FeCl_3$
ii) 	(M) Effervescence with $NaHCO_3$
iii) 	(N) Yellow precipitate with NaOH and I_2
iv) 	(O) Yellow oil with $NaNO_2$, HCl at 0 ^oC (P) Heating with NaOH gives out a gas that turns moist turmeric paper brown

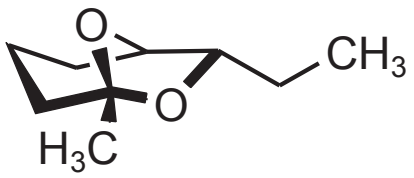
(A) i)-N ii)-L iii)-O iv)-M
(B) i)-O ii)-N iii)-L iv)- P
(C) i)-P ii)-O iii)-L iv)-M
(D) i)-P ii)-N iii)-O iv)-L

22. While doing titration, a student recorded a burette reading of 10.0 mL for the neutralization of 10.0 mL $NaHC_2O_4$ (aq) with 0.1 M NaOH (aq). In a separate experiment, 10.0 mL of this $NaHC_2O_4$ (aq) solution could be completely oxidized by 10.0 mL of $KMnO_4$ in an acidic medium.

What would be the molarity of $KMnO_4$ used by this student?

(A) 0.02 M
(B) 0.04 M
(C) 0.1 M
(D) 0.2 M

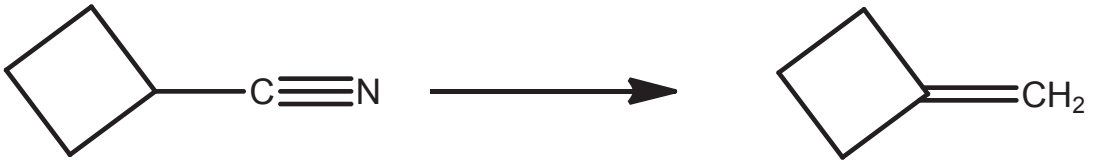
23. Pheromones are chemicals that animals produce for social response. The structure of brevicomin, a pheromone, is shown below. The open chain ketodiol that would form brevicomin is



Brevicomin

(A) 7,8-dihydroxynonan-3-one
(B) 6,7-dihydroxynonan-3-one
(C) 7,8-dihydroxynonan-2-one
(D) 6,7-dihydroxynonan-2-one

24. The best reagents and conditions to accomplish the following conversion is



(A) (i) $LiAlH_4$ in ether, (ii) 3 moles of CH_3I followed by heating with AgOH
(B) (i) $LiAlH_4$ in ether; (ii) P_2O_5 and heat
(C) (i) 20 % H_2SO_4 & heat, (ii) P_2O_5 and heat
(D) H_2 and Lindlar catalyst

PART A-2 · ONE OR MORE CORRECT OPTIONS

Q. 25 – 32 · 6 marks only if all correct options are marked · no negative marking

25. The correct statement/s among the following is/are

(A) Intermolecular forces in n-heptane are stronger than those in 2-methylheptane
(B) Boiling point of 2,2-dimethylpentane is higher than that of 2, 2-dimethylbutane
(C) Both hydrogen bonding and van der Waals forces exist between molecules of 2-methylbutan-2-ol
(D) In 2,2-dimethylbutane, 1°, 2° and 3° types of carbon atoms are present

26. Which of the following aqueous solution/s will have a pH value between 4.0 and 5.0 at 25 °C?

(A) 0.01 M solution of benzoic acid ($K_a = 6.6 \times 10^{-5}$ at 25 °C)
(B) 0.02 mol benzoic acid and 0.05 mol sodium benzoate dissolved in appropriate amount of water to make a solution of 1L
(C) A mixture of 999 mL water and 1mL 0.2 M HCl
(D) 499 mL of 0.01M NaOH and 501 mL of 0.01 M HCl mixed together

27. The energy required to remove an electron from a gaseous species 'X' to form 'X^+' is known as first ionization energy (IE) of X. The energy required to remove an electron from a gaseous species 'X^+' to form 'X^{++}' is called the second IE of X. Similarly, the energy required to remove an electron from a gaseous species X^- to form X is called the IE of X^-.

Identify the correct statement/s from the following

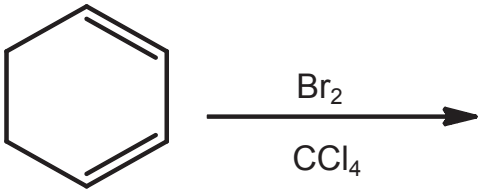
(A) The second IE of the He atom is *four times* that of the (first) IE of the H atom.

(B) The first IEs of F, Ne and Na atoms follow the order IE(Na) < IE(Ne) < IE(F)

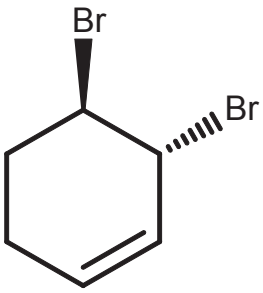
(C) The second IE of the H^- ion is much less than the (first) IE of the H atom.

(D) The IEs of Li, Na and K atoms follow the order IE(K) < IE(Na) < IE (Li)

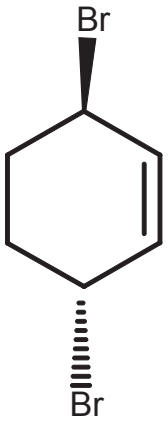
28. The product/s formed in the following reaction is/are



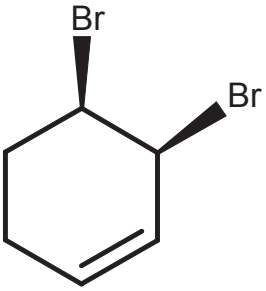
(A)



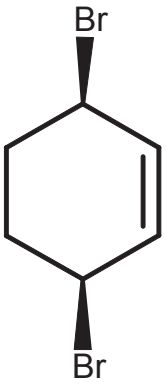
(B)



(C)



(D)



29. Which of the following option/s is /are correct?

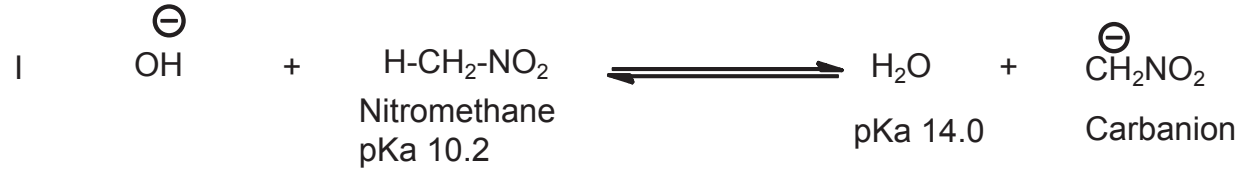
(A) C_2 is paramagnetic

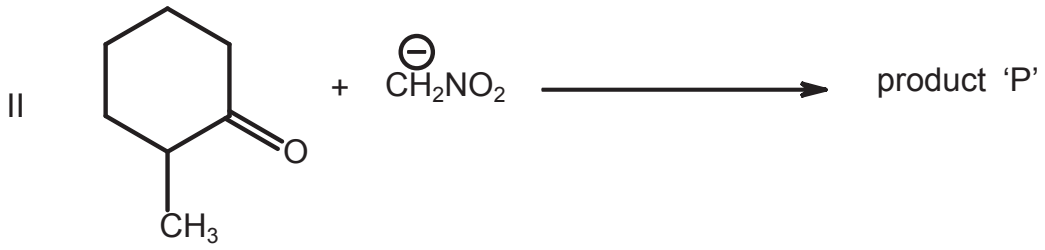
(B) He_2^+ has the same energy as that of two isolated He atoms

(C) S_2 is paramagnetic and S_2^{2-} is diamagnetic

(D) N_2^+ and N_2^- have the same bond order

30. Nitromethane undergoes an aldol type reaction with a racemic mixture of 2-methylcyclohexanone in presence of aqueous NaOH in two steps (I, II) to give the product '**P'**. The statement/s *NOT* correct among the following is/are





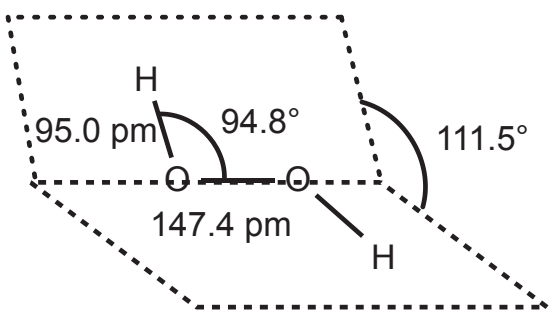
(A) The equilibrium in step I will be more towards the right as water is a stronger acid than nitromethane

(B) The carbanion formed in reaction I can be stabilized due to resonance

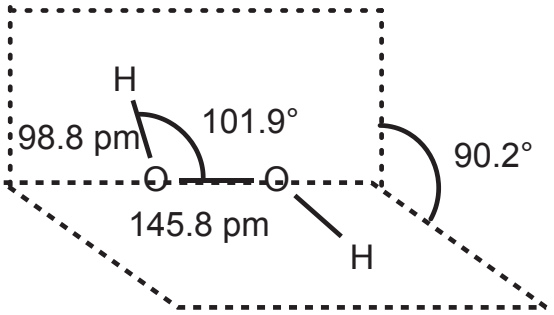
(C) The product formed will be a mixture of four stereoisomers in the form of two pairs of enantiomers

(D) The mixture of products formed can be readily dehydrated to give a single product

31. The structures of hydrogen peroxide (H_2O_2) in the solid and gaseous states are given below. H_2O_2 (l) is slightly more viscous than H_2O (l). The correct option/s among the following is/are



Gas



Solid (Crystal)

(A) Both O atoms are near enough to cause repulsion between the electron lone pairs thus making the O-O bond susceptible for cleavage

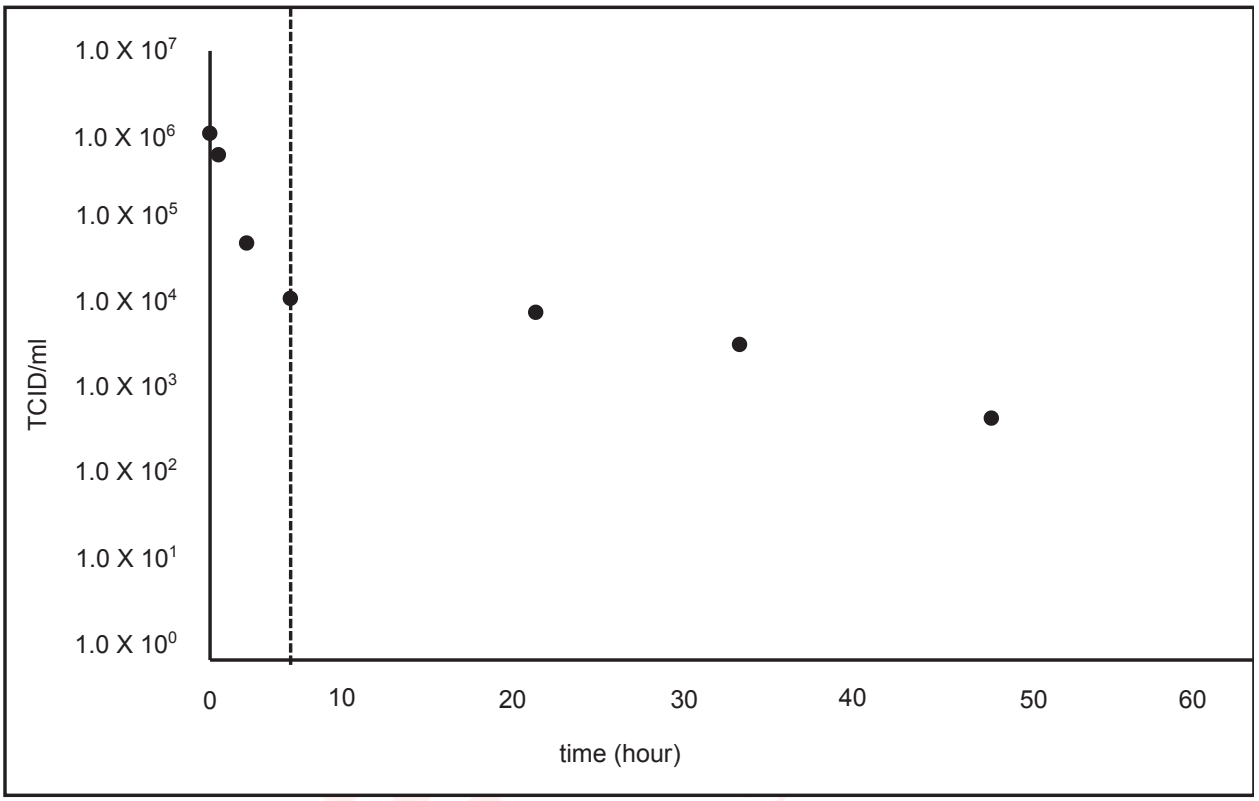
(B) The strong intermolecular H-bonding along with restricted rotation present in the liquid state of H_2O_2 make it more viscous than H_2O (l)

(C) The molecule gets twisted to minimize the repulsion between the lone pair and bond pair of electrons

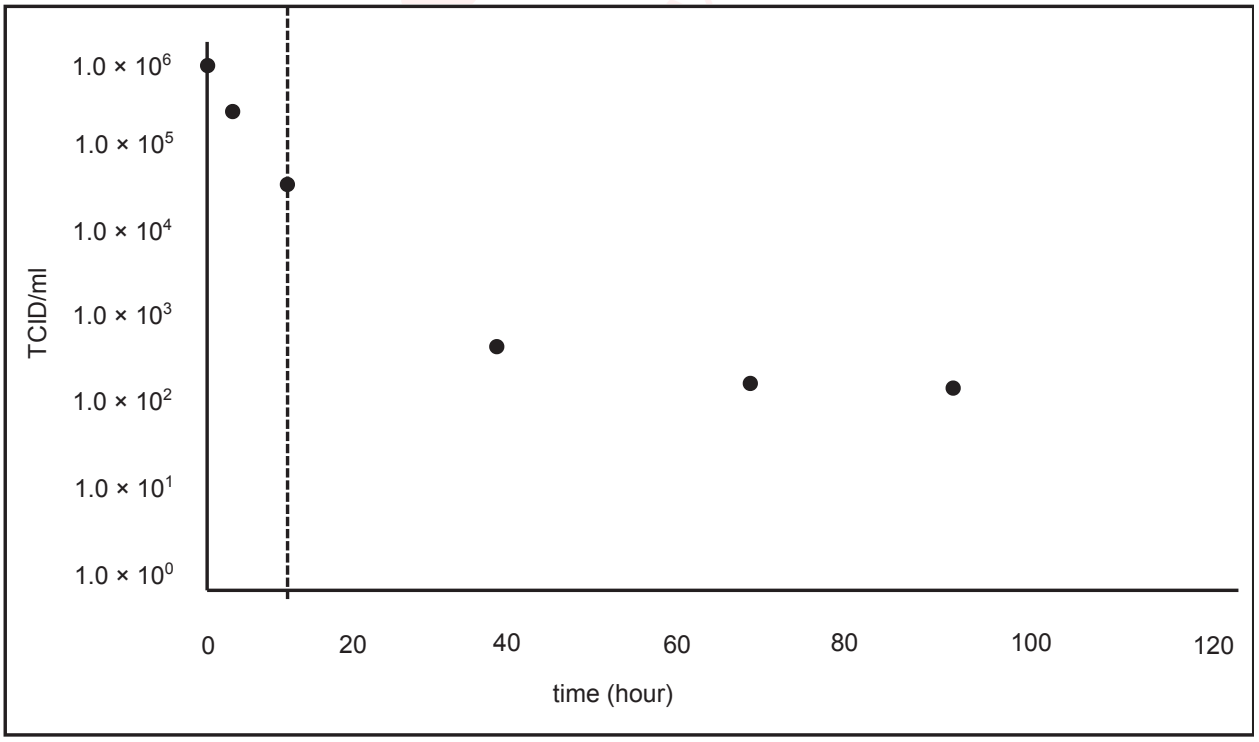
(D) The difference in the dihedral angles in the solid and gaseous states is a consequence of hydrogen bonding between the molecules

32. Viruses are nonliving complex chemical entities. They undergo inactivation and hence lose the ability to infect a host, with time. Concentration (expressed as 'median tissue culture infectious dose', TCID/ml, a unit used in expressing virus concentrations) vs. time plots of a corona virus on the surfaces of a paper currency note and a plastic currency note are shown below. Both these plots have two separate regions (shown by vertical lines in the plots), indicating two time zones.

I. Paper currency note



II. Plastic currency note



The correct option/s among the following is/are

(A) Inactivation of the virus follows zero order kinetics in 1st zone and first order kinetics in 2nd zone

(B) The rate of inactivation is independent of the surface material

(C) The virus reacts with different chemical entities/substances in 1st zone and 2nd zone

(D) On both the surfaces, at least 95 % of the virus is inactivated within 10 h

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Given in the question paper

IUPAC Periodic Table of the Elements

1																	18
1 **H** hydrogen 1.008 [1.0078, 1.0082]	2		Key: atomic number **Symbol** name conventional atomic weight standard atomic weight									13	14	15	16	17	2 **He** helium 4.0026
3 **Li** lithium 6.94 [6.938, 6.997]	4 **Be** beryllium 9.0122											5 **B** boron 10.81 [10.806, 10.821]	6 **C** carbon 12.011 [12.009, 12.012]	7 **N** nitrogen 14.007 [14.006, 14.008]	8 **O** oxygen 15.999 [15.999, 16.000]	9 **F** fluorine 18.998	10 **Ne** neon 20.180
11 **Na** sodium 22.990	12 **Mg** magnesium 24.305 [24.304, 24.307]	3	4	5	6	7	8	9	10	11	12	13 **Al** aluminium 26.982	14 **Si** silicon 28.085 [28.084, 28.086]	15 **P** phosphorus 30.974	16 **S** sulfur 32.06 [32.059, 32.076]	17 **Cl** chlorine 35.45 [35.446, 35.457]	18 **Ar** argon 39.95 [39.792, 39.963]
19 **K** potassium 39.098	20 **Ca** calcium 40.078(4)	21 **Sc** scandium 44.956	22 **Ti** titanium 47.867	23 **V** vanadium 50.942	24 **Cr** chromium 51.996	25 **Mn** manganese 54.938	26 **Fe** iron 55.845(2)	27 **Co** cobalt 58.933	28 **Ni** nickel 58.693	29 **Cu** copper 63.546(3)	30 **Zn** zinc 65.38(2)	31 **Ga** gallium 69.723	32 **Ge** germanium 72.630(8)	33 **As** arsenic 74.922	34 **Se** selenium 78.971(8)	35 **Br** bromine 79.904 [79.901, 79.907]	36 **Kr** krypton 83.798(2)
37 **Rb** rubidium 85.468	38 **Sr** strontium 87.62	39 **Y** yttrium 88.906	40 **Zr** zirconium 91.224(2)	41 **Nb** niobium 92.906	42 **Mo** molybdenum 95.95	43 **Tc** technetium	44 **Ru** ruthenium 101.07(2)	45 **Rh** rhodium 102.91	46 **Pd** palladium 106.42	47 **Ag** silver 107.87	48 **Cd** cadmium 112.41	49 **In** indium 114.82	50 **Sn** tin 118.71	51 **Sb** antimony 121.76	52 **Te** tellurium 127.60(3)	53 **I** iodine 126.90	54 **Xe** xenon 131.29
55 **Cs** caesium 132.91	56 **Ba** barium 137.33	57-71 lanthanoids	72 **Hf** hafnium 178.49(2)	73 **Ta** tantalum 180.95	74 **W** tungsten 183.84	75 **Re** rhenium 186.21	76 **Os** osmium 190.23(3)	77 **Ir** iridium 192.22	78 **Pt** platinum 195.08	79 **Au** gold 196.97	80 **Hg** mercury 200.59	81 **Tl** thallium 204.38 [204.38, 204.39]	82 **Pb** lead 207.2	83 **Bi** bismuth 208.98	84 **Po** polonium	85 **At** astatine	86 **Rn** radon
87 **Fr** francium	88 **Ra** radium	89-103 actinoids	104 **Rf** rutherfordium	105 **Db** dubnium	106 **Sg** seaborgium	107 **Bh** bohrium	108 **Hs** hassium	109 **Mt** meitnerium	110 **Ds** darmstadtium	111 **Rg** roentgenium	112 **Cn** copernicium	113 **Nh** nihonium	114 **Fl** flerovium	115 **Mc** moscovium	116 **Lv** livermorium	117 **Ts** tennessine	118 **Og** oganesson

57 **La** lanthanum 138.91	58 **Ce** cerium 140.12	59 **Pr** praseodymium 140.91	60 **Nd** neodymium 144.24	61 **Pm** promethium	62 **Sm** samarium 150.36(2)	63 **Eu** europium 151.96	64 **Gd** gadolinium 157.25(3)	65 **Tb** terbium 158.93	66 **Dy** dysprosium 162.50	67 **Ho** holmium 164.93	68 **Er** erbium 167.26	69 **Tm** thulium 168.93	70 **Yb** ytterbium 173.05	71 **Lu** lutetium 174.97
89 **Ac** actinium	90 **Th** thorium 232.04	91 **Pa** protactinium 231.04	92 **U** uranium 238.03	93 **Np** neptunium	94 **Pu** plutonium	95 **Am** americium	96 **Cm** curium	97 **Bk** berkelium	98 **Cf** californium	99 **Es** einsteinium	100 **Fm** fermium	101 **Md** mendelevium	102 **No** nobelium	103 **Lr** lawrencium

INTERNATIONAL UNION OF PURE AND APPLIED CHEMISTRY

For notes and updates to this table, see www.iupac.org. This version is dated 1 December 2018.
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United Nations Educational, Scientific and Cultural Organization

2019 IYPT International Year of the Periodic Table of Chemical Elements

IUPAC

ANSWER KEY

IOQC 2020-21 · PAPER CODE 31

32
QUESTIONS · 120 MARKS

IAPT final answer key for paper code 31. ★ marks an answer given as per Zigyan.

1	(A)	2	(C)	3	(B)	4	(D)	5	(C)
6	(B)	7	(B)	8	(D)	9	(B)	10	(B)
11	(A)	12	(A)	13	(B)	14	(A)	15	(C)
16	(A)	17	(C)	18	(A)★	19	(D)	20	(C)
21	(D)	22	(B)	23	(D)	24	(A)	25	(B), (C)
26	(B), (D)	27	(A), (D)	28	(A), (B), (D)	29	(C), (D)	30	(A), (D)
31	(A), (B), (C), (D)★	32	(C), (D)★						

★ Answers as per Zigyan

- **Q18** — As per Zigyan: (A); IAPT key (A).
- **Q31** — As per Zigyan: (A), (B), (C), (D); IAPT key (A), (B), (C), (D).
- **Q32** — As per Zigyan: (C), (D); IAPT key (C), (D).

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OLYMPIADS · 2026

Olympiad results



TOP ACHIEVEMENT · CHEMISTRY

Qualified the Chemistry Olympiad twice, and selected for the HBCSE camp.

Saksham Bishnoi reached the national stage of the Chemistry Olympiad in two consecutive years and was selected for the Orientation-cum-Selection Camp at HBCSE, Mumbai. 4 Years Classroom - the same student who took AIR 168 in JEE Advanced 2026.

Qualified twice | HBCSE selection camp | Only student from Faridabad

Above MAS means above the **Minimum Admissible Score**: clearing it puts a student in the national band for their subject. It is a real threshold, not a participation marker.

CHEMISTRY · ABOVE MAS



Sulokya Bibhuty
Class XII
3 Years Classroom
Above MAS



Anay Gupta
Class XI
3 Years Classroom
Above MAS



Akshat Khator
Class XI
3 Years Classroom
Above MAS



Harshit Goyal
Class XII
4 Years Classroom
Above MAS



Hardik Sharma
Class XI
3 Years Classroom
Above MAS



Aarush Kumar
Class XII
1 Year Classroom
Above MAS

PHYSICS · ASTRONOMY · MATHEMATICS

PHYSICS



Sulokya Bibhuty
Class XII
3 Years Classroom
Above MAS

PHYSICS



Saksham Bishnoi
Class XII
4 Years Classroom
Above MAS

ASTRONOMY



Sulokya Bibhuty
Class XI
3 Years Classroom
Above MAS

MATHEMATICS · IOQM



Priyanshi
Class XI
4 Years Classroom
IOQM Qualifier

OLYMPIADS WE PREPARE STUDENTS FOR

NSEP	NSEC	NSEB	NSEA	IOQM	NSEJS
Physics	Chemistry	Biology	Astronomy	Mathematics	Class X and below

The olympiad pathway starts in Class 8.

NSEJS is open to Class 10 and below - the habit of staying with one hard problem is the one JEE Advanced rewards.

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RESULTS · 2021 – 2026

Our toppers, year by year

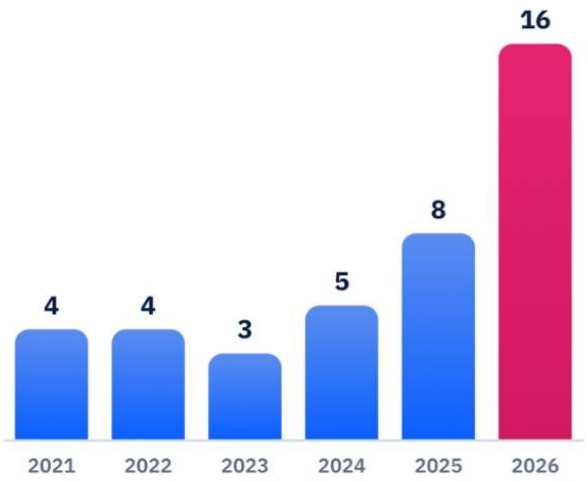
40
JEE Advanced selections in 6 years

AIR 168
best rank — Saksham Bishnoi, 2026

16 of 35
students selected, JEE Advanced 2026

14 of 35
above 98 percentile, JEE Main 2026

JEE ADVANCED SELECTIONS PER YEAR



BEST RANK EACH YEAR

YEAR	JEE ADVANCED	JEE MAIN
2026	**Saksham Bishnoi** AIR 168 · 4 Years Classroom	**Saksham Bishnoi** AIR 811 · 99.95 %ile
2025	**Chakshu** AIR 4,233 · 2 Years Classroom	**Chakshu** AIR 8,538 · 99.44 %ile
2024	**Harshit** AIR 777 · 2 Years Classroom	**Harshit** AIR 1,473 · 99.91 %ile
2023	**Aditya Jain** AIR 2,373 · 2 Years Classroom	**Aditya Jain** AIR 2,503 · 99.79 %ile
2022	**Rachit Sharma** AIR 698 · 1 Year Classroom	**Rachit Sharma** AIR 1,016 · 99.90 %ile
2021	**Sparsh Mittal** AIR 2,237 · 1 Year Classroom	**Sparsh Mittal** AIR 1,562

JEE ADVANCED 2026 — THIRTEEN NAMED SELECTIONS



Saksham Bishnoi
AIR 168
4 Years Classroom



Sulokya Bibhuty
AIR 1,315
3 Years Classroom



Aarush
AIR 2,699
1 Year Classroom



Manishka
AIR 2,869
4 Years Classroom



Oojit Kumar
AIR 5,765
4 Years Classroom



Harshit Goyal
AIR 6,407
4 Years Classroom



Pranab Sharma
AIR 6,750
3 Years Classroom



Aditya Dhar
AIR 15,183
3 Years Classroom



Arpit Anshuman
AIR 19,308
4 Years Classroom



Lokesh
AIR 22,857
3 Years Classroom



Arjun
AIR 24,674
3 Years Classroom



Sanyam Gulyani
AIR 25,123
1 Year Classroom



Saket
AIR 27,628
3 Years Classroom

Sixteen students of the 2026 batch were selected in JEE Advanced; the thirteen above are named with their All India Rank.

JEE MAIN 2026 — TOP SIX, WITH PERCENTILE



Saksham Bishnoi
AIR 811
99.95 %ile
4 Years Classroom



Sulokya Bibhuty
AIR 1,095
99.94 %ile
3 Years Classroom



Harshit Goyal
AIR 2,662
99.83 %ile
4 Years Classroom



Manishka
AIR 4,817
99.70 %ile
4 Years Classroom



Oojit Kumar
AIR 5,040
99.69 %ile
4 Years Classroom



Aarush
AIR 5,276
99.67 %ile
1 Year Classroom

14 of the 35 students in the 2026 batch scored above the 98th percentile in JEE Main.

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